

Chapter 11 Chemical Reactions Answers

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FSc Chemistry Book1, CH 11, LEC 8: Physical Methods for Rates of Reactions **Chapter 11 Chemical Reactions**

Answers

Chemistry chapter 11 Chemical reactions answer key. coefficient. a whole number that appears before a formula in an equation. spectator ion. a particle not directly involved in a chemical reaction. combustion reaction. a reaction in which oxygen reacts with another substance, often producing light or heat. reactant.

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Section 11.1 Assessment. Describe the steps in writing a balanced chemical equation. Write the skeleton equation for the following reactions: Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur dioxide gas. Iron metal and chlorine gas react to form solid iron(III)chloride. $\text{CuS (s)} + \text{O}_2\text{(g)} \rightarrow \text{Cu (s)} + \text{SO}_2\text{(g)}$ D

Chapter 11: Chemical Reactions

Chapter 11 Chemical Reactions Test. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. tagreen348. Terms in this set (51) (aq) dissolved in water (s) solid (l) liquid. reactants. the starting substances of a chemical reaction. products. the ending substances of a chemical reaction.

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330 Chapter 11 11.2 Types of Chemical Reactions Often charcoal briquettes provide the heat for barbeque grills through the burning of carbon. Have you ever felt the heat and smelled the smoke coming from a burning charcoal grill? The heat and smoke are the products of a combustion reaction. Combustion is one of the ?ve general types of chemical

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Write correct formulas for the products in these double replacement reactions. 1) $\text{H}_3\text{PO}_4 + 2\text{K}_2\text{CO}_3 + \text{BaCl}_2 \rightarrow \text{Ba}_3(\text{PO}_4)_2 + 6\text{KCl}$ 2) $\text{K}_2\text{CO}_3 + \text{BaCl}_2 \rightarrow \text{BaCO}_3 + 2\text{KCl}$ 3) $\text{Al}_2(\text{SO}_4)_3 + 3\text{Ca}(\text{OH})_2 \rightarrow 3\text{CaSO}_4 + 2\text{Al}(\text{OH})_3$ 4) $\text{Al}_2(\text{SO}_4)_3 + 3\text{Ca}(\text{OH})_2 \rightarrow 3\text{CaSO}_4 + 2\text{Al}(\text{OH})_3$ 5) $\text{AgNO}_3 + \text{KCl} \rightarrow \text{AgCl} + \text{KNO}_3$ 6) $\text{H}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow \text{BaSO}_4 + 2\text{HCl}$ 7) $\text{BCl}_3 + 3\text{H}_2\text{O} \rightarrow \text{H}_3\text{BO}_3 + 3\text{HCl}$ 8) $\text{H}_2\text{SO}_4 + \text{Ca}(\text{OH})_2 \rightarrow \text{CaSO}_4 + 2\text{H}_2\text{O}$ 9) $\text{Ag}_2\text{CO}_3 \rightarrow 2\text{Ag}_2\text{O} + \text{CO}_2$ 10) $\text{Ag}_2\text{CO}_3 \rightarrow 2\text{Ag}_2\text{O} + \text{CO}_2$ 11) $\text{K}_2\text{Cr}_2\text{O}_7 \rightarrow 2\text{K}_2\text{O} + 2\text{Cr}_2\text{O}_3 + 7\text{O}_2$

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11.2 Types of Chemical Reactions> 29 Whether one metal will displace another metal from a compound depends upon the relative reactivities of the 11.2 Types of Chemical Reactions> Section 11.1 –Describing Chemical Reactions
The reactants are written on the left and the products on the right. The arrow that separates them is called yield. Chapter 11: Chemical Reactions

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11.2 Types of Chemical Reactions> 30 A halogen can also replace another halogen from a compound. •The activity of halogens decreases as you go down Group 7A of the periodic table—fluorine, chlorine, bromine, and iodine. •Bromine is more active than iodine, so this reaction occurs: $\text{Br}_2(\text{aq}) + 2\text{NaI}(\text{aq}) \rightarrow 2\text{NaBr}(\text{aq}) + \text{I}_2(\text{aq})$

11.2 Types of Chemical Reactions>

Most chemical reactions can be classified into one or more of five basic types: acid–base reactions: A chemical reaction that has the general form $\text{AB} + \text{C} \rightarrow \text{AC} + \text{B}$ or $\text{AB} + \text{CD} \rightarrow \text{AD} + \text{CB}$., condensation reactions: A chemical reaction that has the general form $\text{A} + \text{B} \rightarrow \text{AB}$. Condensation reactions are the reverse of cleavage reactions.

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SECTION 11.1 DESCRIBING CHEMICAL REACTIONS
(pages 321 ...

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As shown in Figure 11.3.1, applying a small amount of heat to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the ammonium dichromate volcano. Heat, light, and gas are produced as a large pile of fluffy green chromium(III) oxide forms. We can describe this reaction with a chemical equation An expression that gives the identities and quantities of the ...

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